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CRC Press is pleased to introduce the new edition of Commonly Asked Questions in Thermodynamics, an indispensable resource for those in modern science and engineering disciplines from molecular science, engineering and biotechnology to astrophysics. Fully updated throughout, this edition features two new chapters focused on energy utilization and biological systems. This edition begins by setting out the fundamentals of thermodynamics, including its basic laws and overarching principles. It provides explanations of those principles in an organized manner, using questions that arise frequently from undergraduates in the classroom as the stimulus. These early chapters explore the language of thermodynamics; the first and second laws; statistical mechanical theory; measurement of thermodynamic quantities and their relationships; phase behavior in single and multicomponent systems; electrochemistry; and chemical and biochemical reaction equilibria. The later chapters explore applications of these fundamentals to a diverse set of subjects including power generation (with and without fossil fuels) for transport, industrial and domestic use; heating; decarbonization technologies; energy storage; refrigeration; environmental pollution; and biotechnology. Data sources for the properties needed to complete thermodynamic evaluations of many processes are included. The text is designed for readers to dip into to find an answer to a specific question where thermodynamics can provide some, if not all, of the answers, whether in the context of an undergraduate course or not. Thus its readership extends beyond conventional technical undergraduates to practicing engineers and also to the interested lay person who seeks to understand the discourse that surrounds the choice of particular technological solutions to current and future energy and material production problems.

LIPIDAT is a convenient compilation of thermodynamic data and bibliographic information on lipids. Over 11,000 records in 15 information fields are provided. The book presents tabulations of all known mesomorphic and polymorphic phase transition types, temperatures, and enthalpies for synthetic and biologically derived lipids in dry, partially hydrated, and fully hydrated states. It also includes the effect of pH, protein, drugs, salt, and metal ion concentration on these thermodynamic values. Methods used in making the measurements and the experimental conditions are reported. Bibliographic information includes a complete literature reference and list of authors. The book will be an indispensable reference for biophysicists, chemical engineers, pharmaceutical and cosmetic researchers, dermatologists, nutritionists, biochemists, physiologists, food scientists, and fats and oils chemists.

Thermodynamics sets fundamental laws for all physical processes and is central to driving and maintaining planetary dynamics. But how do Earth system processes perform work, where do they derive energy from, and what are the limits? This accessible book describes how the laws of thermodynamics apply to Earth system processes, from solar radiation to motion, geochemical cycling and biotic activity. It presents a novel view of the thermodynamic Earth system explaining how it functions and evolves, how different forms of disequilibrium are being maintained, and how evolutionary trends can be interpreted as thermodynamic trends. It also offers an original perspective on human activity, formulating this in terms of a thermodynamic, Earth system process. This book uses simple conceptual models and basic mathematical treatments to illustrate the application of thermodynamics to Earth system processes, making it ideal for researchers and graduate students across a range of Earth and environmental science disciplines.

The study of phase transitions is among the most fascinating fields in physics. Originally limited to transition phenomena in equilibrium systems, this field has outgrown its classical confines during the last two decades. The behavior of far from equilibrium systems has received more and more attention and has been an extremely active and productive subject of research for physicists, chemists and biologists. Their studies have brought about a more unified vision of the laws which govern self-organization processes of physico-chemical and biological systems. A major achievement has been the extension of the notion of phase transition to instabilities which occur only in open nonlinear systems. The notion of phase transition has been proven fruitful in application to nonequilibrium instabilities known for about eight decades, like certain hydrodynamic instabilities, as well as in the case of the more recently discovered instabilities in quantum optical systems such as the laser, in chemical systems such as the Belousov-Zhabotinskii reaction and in biological systems. Even outside the realm of natural sciences, this notion is now used in economics and sociology. In this monograph we show that the notion of phase transition can be extend ed even further. It apphes also to a new class of transition phenomena which occur only in nonequilibrium systems subjected to a randomly fluctuating environment.

A quantitative measure of the accuracy of the rate coefficients and the excess energies is a desirable goal of this analysis. There are two major sources of uncertainties: The atomic and molecular data and the solar irradiance. The cross sections and branching ratios used in this analysis come from many different sources; many of them without any error indications. For this reason, we must confine ourselves to a qualitative indication of the reliability of the results. Specifically we give a quality scale in Table II for the data of each mother molecule, A indicating the highest quality of atomic and molecular data and F the lowest quality. The letter B typically means that the threshold is uncertain. For most molecules the cross section at threshold is very small and the rate coefficient for these molecules is therefore not influenced by this uncertainty. For atomic species the cross section is usually large near threshold, but for these species the threshold is known quite accurately. The letter B, therefore, indicates that the rate coefficient is most likely quite accurate, but the excess energy is less accurately known. The letter C usually means that the branching ratios are not well known. This means that the total rate coefficient is very good, but the rate coefficients and the excess energies for the individual branches are less accurate.

Metrology and its applications e.g. in chemical or food analysis or in environmental monitoring are entering our daily life. This book provides a basic overview over the relevant metrological concepts like traceability, ISO uncertainties or cause-and-effect diagrams. The applications described in great detail range from progression-of-error type evaluation of the measurement uncertainty budget to complex applications like pH measurement or speciation calculations for aqueous solutions. The consequences of a measurement uncertainty concept for chemical data are outlined for geochemical modeling applied to transport in the subsurface and to nuclear waste disposal. Special sections deal with the deficits of existing thermodynamic data for these applications and with the current position of chemical metrology in respect to other quality assurance measures, e.g. ISO 900x, GLP, European and U.S.-American standards.

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