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reactions answer key. coefficient. a

whole number that appears before a

formula in an equation. spectator ion.

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A particle not directly involved in a chemical reaction. combustion reaction. a reaction in which oxygen reacts with another substance, often producing light or heat. reactant.

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Heating solid copper(II)sulfide in the presence of oxygen gas produces pure copper and sulfur dioxide gas.
Iron metal and chlorine gas react to form solid iron(III)chloride. $\text{CuS (s)} + \text{O}_2\text{(g)} \rightarrow \text{Cu (s)} + \text{SO}_2\text{(g)}$

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Chemical Reactions

by: tagreen348. Terms in this set (51)
(aq) dissolved in water (s) solid (l)
liquid. reactants. the starting
substances of a chemical reaction.
products. the ending substances of a
chemical reaction.

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Additional Questions & Answers.

Question 1. Give an example of a complex chemical reaction. Answer:

‘ Photosynthesis ’ is a complex chemical reaction. Question 2. Give an example of simple chemical reactions. Answer: Rusting of iron and combustion of fuels are simply chemical reactions. Question 3.

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330 Chapter 11 11.2 Types of
Chemical Reactions Often charcoal
briquettes provide the heat for
barbeque grills through the burning
of carbon. Have you ever felt the heat
and smelled the smoke coming from
a burning charcoal grill? The heat and
smoke are the products of a
combustion reaction. Combustion is
one of the five general types of
chemical

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Write correct formulas for the
products in these double
replacement reactions. $\text{H}_3\text{PO}_4 + 2)$
 $\text{K}_2\text{CO}_3 + \text{BaCl}_2 \rightarrow$ 3) $+ \text{Al}_2\text{O}_3$ Co

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Chemical Reactions

5) $\text{AgNO}_3 + \text{KCl} \rightarrow \text{AgCl} + \text{KNO}_3$
Q) $\text{BCl}_3 + \text{H}_2\text{O} \rightarrow \text{H}_3\text{BO}_3 + \text{HCl}$
361) $10) \text{Ag}_2\text{C}_2\text{H}_3\text{O}_2 + \text{K}_2\text{CrO}_4$

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of Chemical Reactions > 29 Whether
one metal will displace another metal

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from a compound depends upon the relative reactivities of the 11.2 Types of Chemical Reactions> Section 11.1 –Describing Chemical Reactions The reactants are written on the left and the products on the right. The arrow that separates them is called yield.

Chapter 11: Chemical Reactions

Chapter 11 1 Chemical Reactions

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11.2 Types of Chemical Reactions> 30

A halogen can also replace another halogen from a compound. •The activity of halogens decreases as you go down Group 7A of the periodic table—fluorine, chlorine, bromine, and iodine. •Bromine is more active than iodine, so this reaction occurs: $\text{Br}_2(\text{aq}) + 2\text{NaI}(\text{aq}) \rightarrow 2\text{NaBr}(\text{aq}) + \text{I}_2(\text{aq})$

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11.2 Types of Chemical Reactions >

Most chemical reactions can be classified into one or more of five basic types: acid–base reactions: A chemical reaction that has the general form $AB + C \rightarrow AC + B$ or $AB + CD \rightarrow AD + CB$., condensation reactions: A chemical reaction that has the general form $A + B \rightarrow AB$. Condensation reactions are the reverse of cleavage reactions.

Chapter 11.6: Types of Chemical Reactions - Chemistry ...

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As shown in Figure 11.3.1, applying a small amount of heat to a pile of orange ammonium dichromate powder results in a vigorous reaction known as the ammonium dichromate volcano. Heat, light, and gas are produced as a large pile of fluffy green chromium(III) oxide forms. We can describe this reaction with a chemical equation An expression that gives the identities and quantities of the ...

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